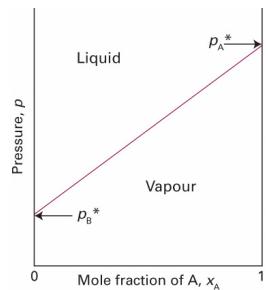


Fazni dijagrami binarnih smjesa

Ako tekućine slijede Raoultov zakon:

$$p = p_A + p_B$$

$$p = p_B^* + (p_A^* - p_B^*) x_A$$

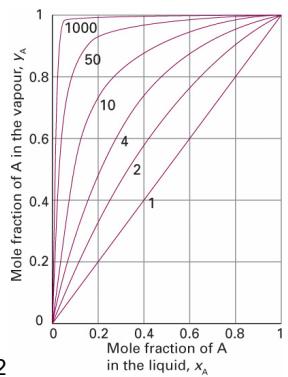


1

Binarne smješe

$$y_A = \frac{x_A p_A^*}{p_B^* + (p_A^* - p_B^*) x_A}$$

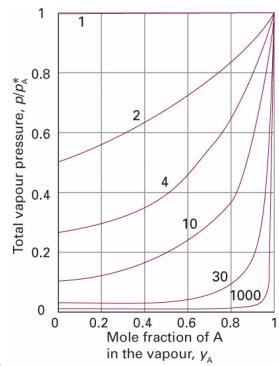
$$y_B = 1 - y_A$$



2

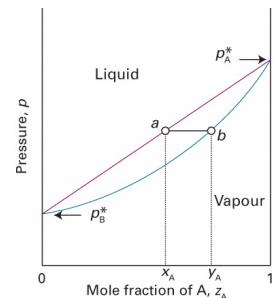
Binarne smješe

$$p = \frac{p_A^* p_B^*}{p_A^* + (p_B^* - p_A^*) y_A}$$



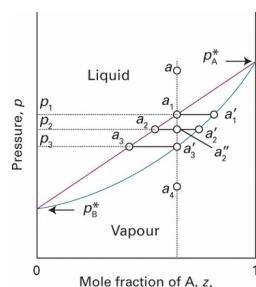
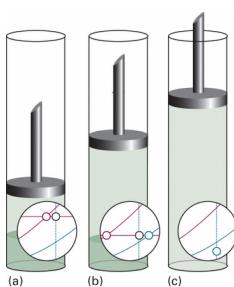
3

Binarne smješe



4

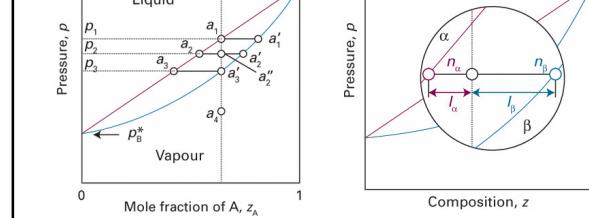
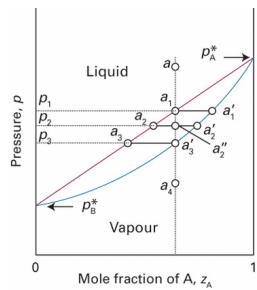
Binarne smješe



5

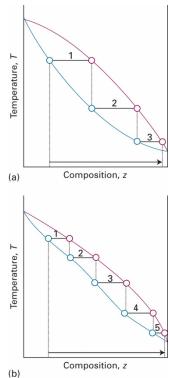
Binarne smješe

$$n_\alpha l_\alpha = n_\beta l_\beta$$



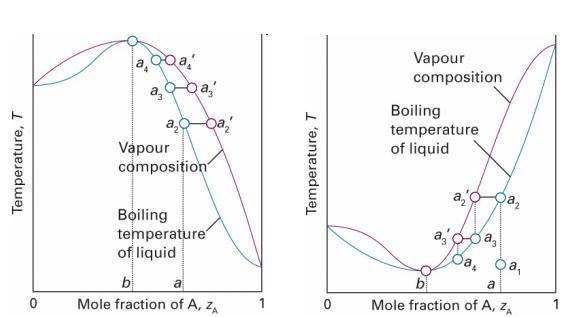
6

Frakcijska destilacija



7

Azeotropi

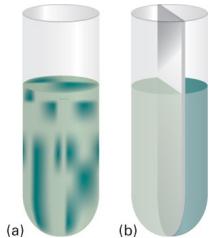


8

Destilacija binarne smjese

Tekućine koje se ne miješaju međusobno

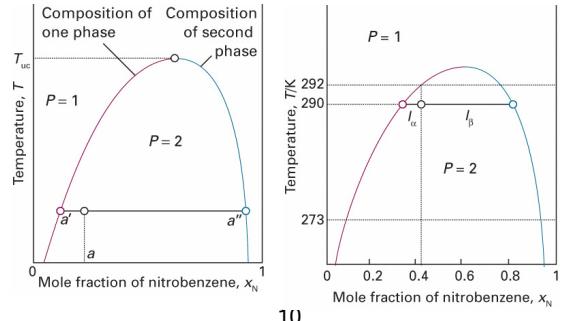
$$p = p_A^* + p_B^*$$



9

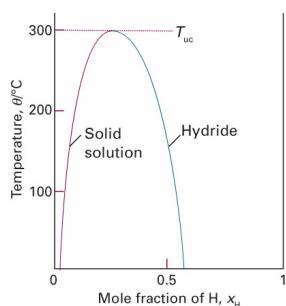
Binarne smjese

Tekućine koje se djelomično miješaju međusobno

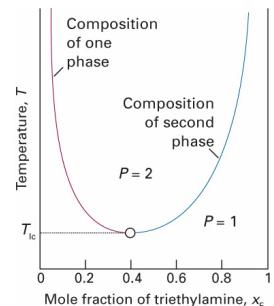


10

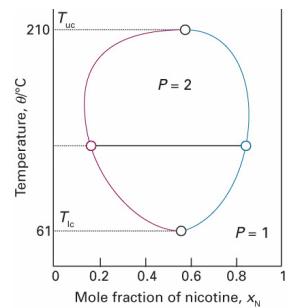
Kritične temperature smjesa



11



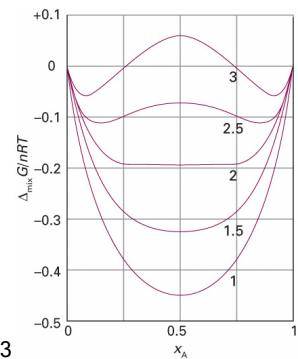
Kritične temperature smjesa



12

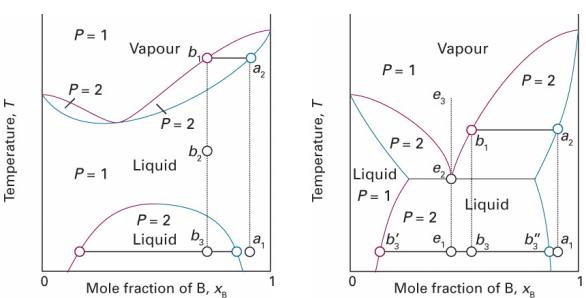
Fazni dijagrami binarnih sistema

$$\Delta G_{\text{mix}} = nRT(x_A \ln x_A + x_B \ln x_B + \xi x_A x_B)$$



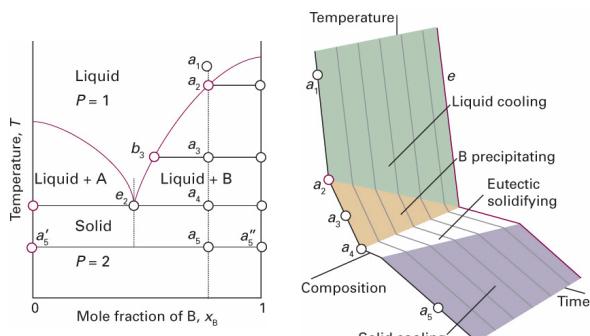
13

Destilacija binarnih smjesa



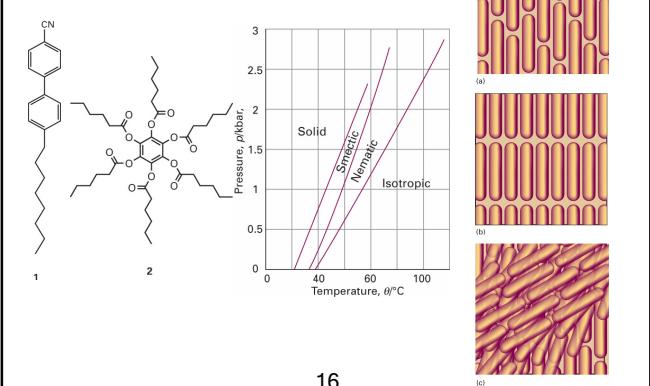
14

Fazni dijagram krutina-tekućina



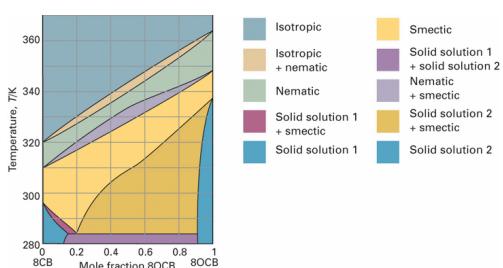
15

Mezofaze



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Mezofaze



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Aktivitet

Aktivitet otapala A:

$$\mu_A = \mu_A^* + RT \ln a_A$$

$$a_A = \gamma_A x_A \quad \gamma_A \rightarrow 1 \quad x_A \rightarrow 1$$

$$\mu_A = \mu_A^* + RT \ln x_A + RT \ln \gamma_A$$

Aktivitet otopljenje tvari B:

$$\mu_B = \mu_B^\circ + RT \ln a_B$$

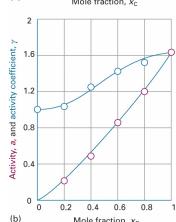
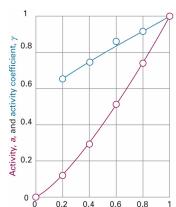
$$a_B = \gamma_B x_B \quad \gamma_B \rightarrow 1 \quad x_B \rightarrow 0$$

$$\mu = \mu^\circ + RT \ln a \quad \gamma_B \rightarrow 1 \quad x_B \rightarrow 0$$

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Aktivitet

$$\mu_A = \mu_A^* + RT \ln a_A$$



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Biološko standardno stanje

$$\mu(H^+) = \mu^\circ(H^+) + RT \ln(H^+)$$

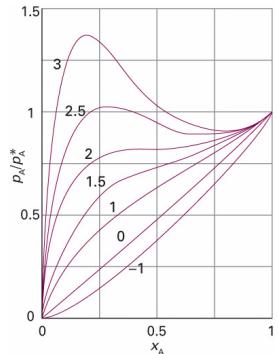
$$\mu^\oplus(H^+) = \mu^\circ(H^+) - RT \ln 10$$

$$RT \ln 10 = 39.96 \text{ kJ mol}^{-1}$$

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Aktivitet

$$\frac{P_A}{P_A^*} = x_A e^{\xi(1-x_A)^2}$$



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